

Alkali metal, any of the six elements of Group 1 (Ia) of the periodic table--lithium, sodium, potassium, rubidium, cesium, and francium. The alkali metals are so called because ...

The alkali metals are all members of group 1 on the periodic table, minus hydrogen. This is the first column of the periodic table. They include lithium, sodium, ...

The table summarizes the important physical and thermodynamic properties of the alkali metals. At atmospheric pressure these metals are all characterized by a body-centred cubic ...

Lithium, chemical element of Group 1 (Ia) in the periodic table, the alkali metal group, lightest of the solid elements. The metal itself--which is soft, white, and lustrous--and several of its ...

The oxides of the alkaline -earth metals are basic (i.e., alkaline, in contrast to acidic). A fairly steady increase in electropositive character is observed in passing from beryllium, the lightest member of ...

Alkaline-earth metal - Properties, Reactivity, Uses: The alkaline-earth elements are highly metallic and are good conductors of electricity. They have a gray-white lustre when freshly cut but tarnish readily ...

They're called alkali metals because when they react with water, they form alkaline solutions - which means basic or high pH solutions. These reactions produce a metal hydroxide and ...

People commonly use the term alkaline for basic solutions, but their meanings are not the same. All alkaline solutions are basic, but not all bases are ...

This page showcases alkali metals, covering their isolation, properties, and reactions. Their high reactivity, strong reducing abilities, and ...

The alkali metals and the alkaline earth metals, as well as the transition metals and the posttransition metals (in their lower oxidation states), form ionic oxides--i.e., compounds that contain the O²⁻ ...

The alkali metals, in Group 1 (Ia), are soft metallic solids with low melting points. The alkaline-earth metals, in Group 2 (IIa), are harder and have higher melting points than the adjacent ...

Alkali, any of the soluble hydroxides of the alkali metals--i.e., lithium, sodium, potassium, rubidium, and cesium. Alkalies are strong bases that turn litmus paper from red to blue; they react with acids to ...

sodium (Na), chemical element of the alkali metal group (Group 1 [Ia]) of the periodic table. Sodium is a very

soft silvery-white metal. Sodium is the most common alkali metal and the sixth most ...

This family of elements is also known as the lithium family after its leading element. The alkali metals are all shiny, soft, highly reactive metals at standard ...

Overview Etymology Common properties of alkalis and bases Difference between alkali and base Alkali salts Alkaline soil Alkali lakes In chemistry, an alkali is a basic salt of an alkali metal or an alkaline earth metal. An alkali can also be defined as a base that dissolves in water. A solution of a soluble base has a pH greater than 7.0. The adjective alkaline, and less often, alkalescent, is commonly used in English as a synonym for basic, especially for bases soluble in water. This broad use of the term is likely to have come about because alkalis were the first bases known to obey the Arrhenius definition of a base, and they are still among th...

What is Alkali? Alkali is a basic, ionic salt of an alkali metal or an alkaline earth metal.

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